

Lect 34

I best time

→ All things H

II More than one quantum

Object: herding cats

I. completed the derivation of the hydrogen atom wavefunctions

Let's reassemble all the pieces!

$$\hat{H}\Psi = E\Psi$$

with $\Psi = \Psi(\vec{r}, t)$

→ $\Psi_n = \psi_n(r, \theta, \phi) e^{-iE_n t / \hbar}$

Then

$$\Psi_n(r, \theta, \phi) = R(r) Y_n(\theta, \phi)$$

leads to

$$Y_l^m(\theta, \phi) = e^{im\phi} \sqrt{\frac{(2l+1)(l-|m|)!}{4\pi(l+|m|)!}}$$

with

$$P_l^m(x) \equiv (1-x^2)^{|m|/2} \left(\frac{d}{dx}\right)^{|m|} P_l(x)$$

and

$$P_l(x) \equiv \frac{1}{2^l l!} \left(\frac{d}{dx}\right)^l (x^2-1)^l$$

For the radial piece

we have found

$$R_{nl}(r) = \frac{1}{r} \rho^{l+1} e^{-\rho} U(\rho)$$

with $\rho \equiv k r$, $K = \frac{\sqrt{-2mE}}{\hbar}$ and

$$U(\rho) = \sum_{j=0}^{j_{\max}} C_j \rho^j$$

$$\text{here } C_{j+1} = \frac{2(j+l+1-n)}{(j+1)(j+2l+2)} C_j$$

and $j_{\max} = n - l - 1$.

Note then that the radial wave

known named function set

$$U(\rho) = L_{n-l-1}^{2l+1}(2\rho)$$

with

$$L_{q-p}^p(x) = (-1)^p \left(\frac{d}{dx}\right)^p L_q(x) \quad \begin{cases} p = 2l+1 \\ q = n+l \end{cases}$$

and $L_q(x) \equiv e^x \left(\frac{d}{dx}\right)^q (e^{-x} x^q)$

the associated and Laguerre polynomials.

functions are labelled $P_{l/s}$ by two quantum numbers n , the principle quant #, and l , the azimuthal quant. #:

$$R(r) = R_{nl}(r)$$

In fact the series (*) with coeff. s is a well

Then, the energies are

$$E_n = - \left[\frac{m}{2\hbar^2} \left(\frac{e^2}{4\pi\epsilon_0} \right)^2 \right] \frac{1}{n^2}$$

$$n = 1, 2, 3, 4, \dots$$

The Bohr radius is

$$a = \frac{\hbar^2}{m} \frac{4\pi\epsilon_0}{e^2} = 0.529 \times 10^{-10} \text{ m}$$

and

$$E_n = - \left[\frac{\hbar^2}{2m} \frac{1}{a^2} \right] \frac{1}{n^2}$$

The wave functions are

$$\psi_{nlm} = \left[\frac{2}{na} \right]^3 \frac{(n-l-1)!}{2n[(n+l)!]^3} e^{-r/na} \left(\frac{r}{na} \right)^l \left[L_{n-l-1}^{2l+1} \left(\frac{2r}{na} \right) \right] Y_l^m(\theta, \phi)$$

Pretty glorious!

Here we denote the combined system state by juxtaposition. The space of all such states is denoted $\mathcal{H}_1 \otimes \mathcal{H}_2$.

What is the total angular momentum of the atom (say the e is in ^{the} ground state)?

II Say you want to combine one spin, say that of an electron, with a second one, e.g. the proton spin. These two states come from different Hilbert spaces, call 'em \mathcal{H}_1 and \mathcal{H}_2 . The combined system has four possibilities $\uparrow\uparrow, \uparrow\downarrow, \downarrow\uparrow, \downarrow\downarrow$

So, if the combined state is $\chi = \chi_1 \chi_2$ then

We can check this using the combined operators

$$S_x = S_x^{(1)} + S_x^{(2)} \quad \begin{matrix} \swarrow \text{acts on } \chi_1 \\ \searrow \text{acts on } \chi_2 \end{matrix}$$

$$\begin{aligned} S_z \chi_1 \chi_2 &= (S_z^{(1)} + S_z^{(2)}) \chi_1 \chi_2 \\ &= (S_z^{(1)} \chi_1) \chi_2 + \chi_1 (S_z^{(2)} \chi_2) \\ &= \hbar m_1 \chi_1 \chi_2 + \hbar m_2 \chi_1 \chi_2 \\ &= \hbar (m_1 + m_2) \chi_1 \chi_2 \end{aligned}$$

- Then
- $\uparrow \uparrow : m=1$
 - $\uparrow \downarrow : m=0$
 - $\downarrow \uparrow : m=0$
 - $\downarrow \downarrow : m=-1$

So, it is this state that we should think of as $s=1, m=0!$

Why are there two $m=0$ states? Cool surprise,

This leads us to collect

$$\left\{ \begin{array}{l}
 |111\rangle = \uparrow\uparrow\uparrow \\
 |110\rangle = \frac{1}{\sqrt{2}}(\uparrow\downarrow + \downarrow\uparrow) \\
 |1-1\rangle = \downarrow\downarrow
 \end{array} \right\} \quad \begin{array}{l}
 s=1 \\
 \text{(Triplet)}
 \end{array}$$

and

$$\left\{ |100\rangle = \frac{1}{\sqrt{2}}(\uparrow\downarrow - \downarrow\uparrow) \right\} \quad \begin{array}{l}
 s=0 \\
 \text{(singlet)}
 \end{array}$$